

## Chapter 14 Chemical Equilibrium

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## Chapter 14. CHEMICAL EQUILIBRIUM

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PLAY. Match. Gravity. Created by. afc5xv. Terms in this set (90) reaction quotient

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(Q) The \_\_\_\_\_ is a measure of the progress of a reaction toward equilibrium.  $Q > K$ . If reaction goes to the left (towards reactants) then Q compared to K is \_\_\_\_\_

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## **literary terms chapter 14 chemistry chemical equilibrium ...**

Chapter 14 Page 1 CHAPTER 14: CHEMICAL EQUILIBRIUM Part One: Describing Chemical Equilibrium A. Basic Concepts. 1. All chemical reactions are, in principle, , i.e., they can go both directions. 2. The symbol means reactions are taking place in forward and reverse directions simultaneously. 3.

## **Chapter 14. Chemical Equilibrium**

Chapter 14. CHEMICAL EQUILIBRIUM 14.1 THE CONCEPT OF EQUILIBRIUM AND THE EQUILIBRIUM CONSTANT Many chemical reactions do not go to completion but instead attain a state of chemical equilibrium. Chemical equilibrium: A state in which the rates of the forward and reverse reactions are equal

## **CHAPTER-14 =====**

Chapter 14 Chemical Equilibrium. Instructor's Resource Materials (Download only) for Chemistry, 7e © 2016 Pearson Education, Inc. John E. McMurry, Robert C. Fay, Jill Robinson. The Equilibrium State. Chemical Equilibrium: The state reached when

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the concentrations of reactants and products remain constant over time.

### **CHAPTER 14: CHEMICAL EQUILIBRIUM**

Chapter 14. Chemical Equilibrium What we will learn: • Concept equilibrium • Equilibrium constant • Writing equilibrium constant expressions • Chemical kinetics and chemical equilibrium • Meaning of equilibrium constant • Factors that affect chemical equilibrium

### **Chapter 14.2: The Equilibrium Constant - Chemistry LibreTexts**

Chapter 14 Learning Outcomes • Describe equilibrium for a chemical reaction in terms of forward and reverse reaction rates. • Describe equilibrium for a chemical reaction in terms of its position. • Explain why forward and reverse reaction rates change so as to obtain equilibrium. • Write the equilibrium expression constant for a ...

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The  $K_p$  for this reaction is  $2.14 \times 10^{-2}$  at about 540 K. Under one set of equilibrium conditions, the partial pressure of ammonia is  $P(\text{NH}_3) = 0.454 \text{ atm}$ ,

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that of hydrogen is  $P(\text{H}_2) = 2.319 \text{ atm}$ , and that of nitrogen is  $P(\text{N}_2) = 0.773 \text{ atm}$ . If an additional 1 atm of hydrogen is added to the reactor to give  $P(\text{H}_2) = 3.319 \text{ atm}$ , how will the system respond? Because the stress is an increase in P ...

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Chapter 14: Chemical Equilibrium. [1] Which is the correct equilibrium constant expression for the following reaction?  $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{Fe}(\text{s}) + 3\text{H}_2\text{O}(\text{g})$

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Position of Chemical Equilibrium ♦the equilibrium position refers to the relative amounts of reactants and products in the system at the point of equilibrium ♦a reaction with an equilibrium position that favors the products:  $[\text{product}] > [\text{reactant}]$  at equilibrium equilibrium lies to the right

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Chapter 14: Chemical Equilibrium Expand/collapse global location Chapter 14.2: The Equilibrium Constant ... the system described in Equation 14.2.19 will reach chemical equilibrium only if a stoichiometric amount of solid carbon or excess solid carbon has been added so that some is still present once the system has reached equilibrium.

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Chapter 14: Chemical Equilibrium. ... Ratio of the concentrations of the products to the concentrations of the reactants at any point during the reaction aside from equilibrium, where each reactant and product in the expression is raised to the power of its stoichiometric coefficient. Commonly denoted by  $Q$ .

### **Chapter 14.5: Factors That Affect Equilibrium - Chemistry ...**

14.4 Expressing the Equilibrium Constant in Terms of Pressure Previously, expressed equilibrium constant w/ concentrations of reactants and products in gaseous reactions the partial pressure of a particular gas is proportional to its concentration \*Reminder\* Partial pressure ( $P_n$ ) = the pressure due to any

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individual component in a gas mixture

### **Chapter 14.1: The Concept of Chemical Equilibrium ...**

Chapter 14: Chemical Equilibrium Q1. A reaction with an equilibrium constant  $K_c = 1.5 \times 10^{21}$  would consist of which of the following at equilibrium: A) approximately equal reactants and products B) some reactants and products with reactants slightly favored C) some reactants and products with products slightly favored D) essentially all reactants

### **dynamic equilibrium Chapter 14: requirements Chemical ...**

chemical equilibrium. a chemical reaction in which the products re-form the original.... a chemical reaction that proceeds to such an extent that at le.... a reversible reaction in which there is no longer any change i.... a state of balance in which the rate of a forward reaction equ.... reversible reaction.

### **Introduction Chapter 14: Chemical Equilibrium - YouTube**

Chemical equilibrium is a dynamic process consisting of forward and reverse reactions that proceed at equal rates. At equilibrium, the composition of the

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system no longer changes with time. The composition of an equilibrium mixture is independent of the direction from which equilibrium is approached.



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