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A sample of a compound contains 3.86 g of sulfur and 4.08 g of oxygen. The mole ratio of sulfur to oxygen in the compound is 1:1. The empirical formula of the compound is SO .

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A sample of a compound contains 3.86 g of sulfur and 4.08

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AnalysisData Analysis 1. A sample of aluminum is placed in a 25-mL graduated cylinder containing 10.0 mL of water. The level of water rises to 18.0 mL. Aluminum has a density of 2.7 g/mL. Calculate the mass of the sample. 2. Saturn is about 1 429 000 km from the Sun.

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CHAPTER 10 SOLUTIONS MANUAL The MoleThe Mole Solutions Manual Chemistry: Matter and Change • Chapter 10 161 Section 10.1 Measuring Matter page 320–324 Practice Problems pages 323–324 1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.
2.50 mol Zn

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Chemistry Matter and Change: Chapter 10. The process by which the atoms of one or more substances are rearranged to form different substances; occurs when there is change in temperature, color, odor, and physical change. The starting substance in a chemical equation.

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