

Study Guide For Chemical Reactions

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Chemical Reactions. The standard representation of a chemical reaction shows an arrow pointing from the reactants to the products: For most chemical reactions in this book, solids are labeled (s), liquids (l), and gases (g). The numerical coefficients in front of the chemical formulas express the moles of each compound or element.

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All atoms present in the reactants are present in the products. Coefficient. A number in front of a chemical formula that indicates how many molecules take part in the reaction. Synthesis reaction. A new compound is formed by the combination of similar reactants. Synthesis reaction. $N + O_2 \rightarrow NO_2$ is what.

Chemical Reactions

STUDY GUIDE Section 9.2 Classifying Chemical Reactions In your textbook, read about synthesis, combustion, decomposition, and replacement reactions. Assume that Q, T, X, and Z are symbols for elements. Match each equation in Column A with the reaction type it represents in Column B. Column A 2. $Q + Z \rightarrow QZ$ Answer the following questions. a. b. c. d.

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must conserve the balance of mass. identify which number indicates the number of atoms of each element in a molecule. the number 6. minimum amount of energy required for reactants to form products. activation energy. substance that lowers energy needed to start a chemical reaction. catalyst.

Bing: Study Guide For Chemical Reactions

Covalent reactions. Atoms can form chemical bonds by sharing electrons. Atoms with more than 3 and less than 6 electrons in the outer shell tend to form chemical bonds by sharing electrons. Consider the reaction between hydrogen and carbon. Hydrogen has 1 electron in the outer shell.

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Chemical reactions. When a chemical reaction occurs, the original substances put together, called reactants, lose their chemical properties and become different substances called products with a different set of chemical properties. Reactions where energy is released are called exothermic reactions. When energy is absorbed, it is called an endothermic reaction.

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During chemical reactions, the molecules of a substance are broken apart and rearranged into new and different molecules. It would be like taking apart the hydrogen and oxygen atoms of water, and...

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CHEMICAL REACTIONS 1. Reactants are on the left side of the equation. They are substances that are changed during a chemical reaction.

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Chapter 11: Chemical Reactions Study Guide □□chemical equation A representation of a chemical reaction with reactants on the left, products on the right, and an arrow separating the two skeleton

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